

# OCR

Oxford Cambridge and RSA

## A Level Chemistry A

H432/02 Synthesis and analytical techniques

**Tuesday 12 June 2018 – Afternoon**

**Time allowed: 2 hours 15 minutes**



**You must have:**

- the Data Sheet for Chemistry A (sent with general stationery)

**You may use:**

- a scientific or graphical calculator



First name

Last name

Centre number

Candidate number

### INSTRUCTIONS

- Use black ink. You may use an HB pencil for graphs and diagrams.
- Complete the boxes above with your name, centre number and candidate number.
- Answer **all** the questions.
- Write your answer to each question in the space provided. If additional space is required, use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.
- Do **not** write in the barcodes.

### INFORMATION

- The total mark for this paper is **100**.
- The marks for each question are shown in brackets [ ].
- Quality of extended responses will be assessed in questions marked with an asterisk (\*).
- This document consists of **32** pages.

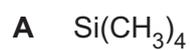
2  
SECTION A

You should spend a maximum of 20 minutes on this section.

Write your answer to each question in the box provided.

Answer all the questions.

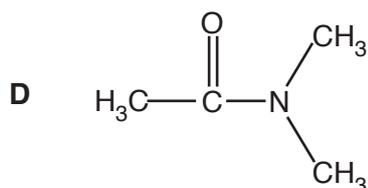
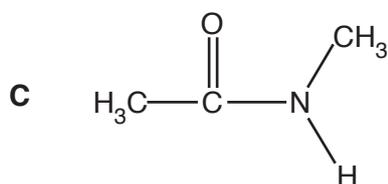
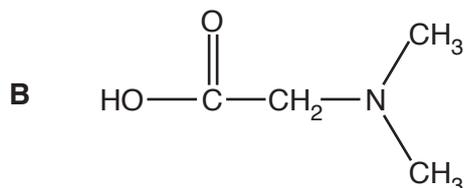
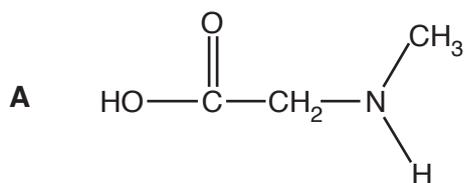
1 Which compound is used as a standard for NMR chemical shift measurements?



Your answer

[1]

2 Which compound is a secondary amide?



Your answer

[1]

3

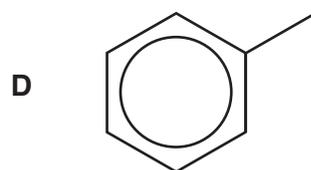
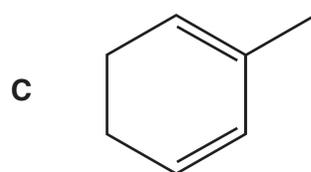
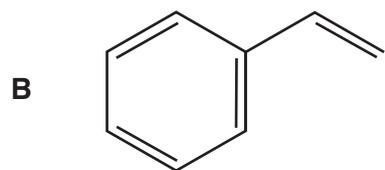
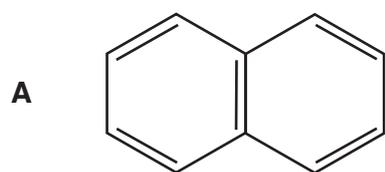
3 Which compound does **not** react with nucleophiles?

- A  $\text{CH}_3\text{CH}_2\text{CHO}$
- B  $\text{CH}_3\text{CHCH}_2$
- C  $\text{CH}_3\text{CH}_2\text{COCH}_3$
- D  $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$

Your answer

[1]

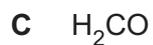
4 Which structure represents an alicyclic compound?



Your answer

[1]

5 Which molecule is **not** planar?



Your answer

[1]

6 What is the number of peaks in the  $^1H$  NMR spectrum of  $HOOCCH_2CHOHCH_2COOH$ ?

A 3

B 4

C 5

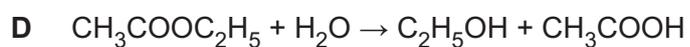
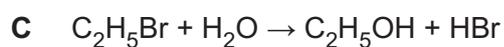
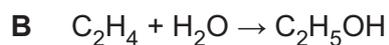
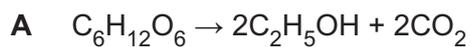
D 6

Your answer

[1]

7 Ethanol can be prepared by different reactions.

Which reaction has the lowest atom economy?



Your answer

[1]

8 The breakdown of ozone is catalysed by NO radicals.

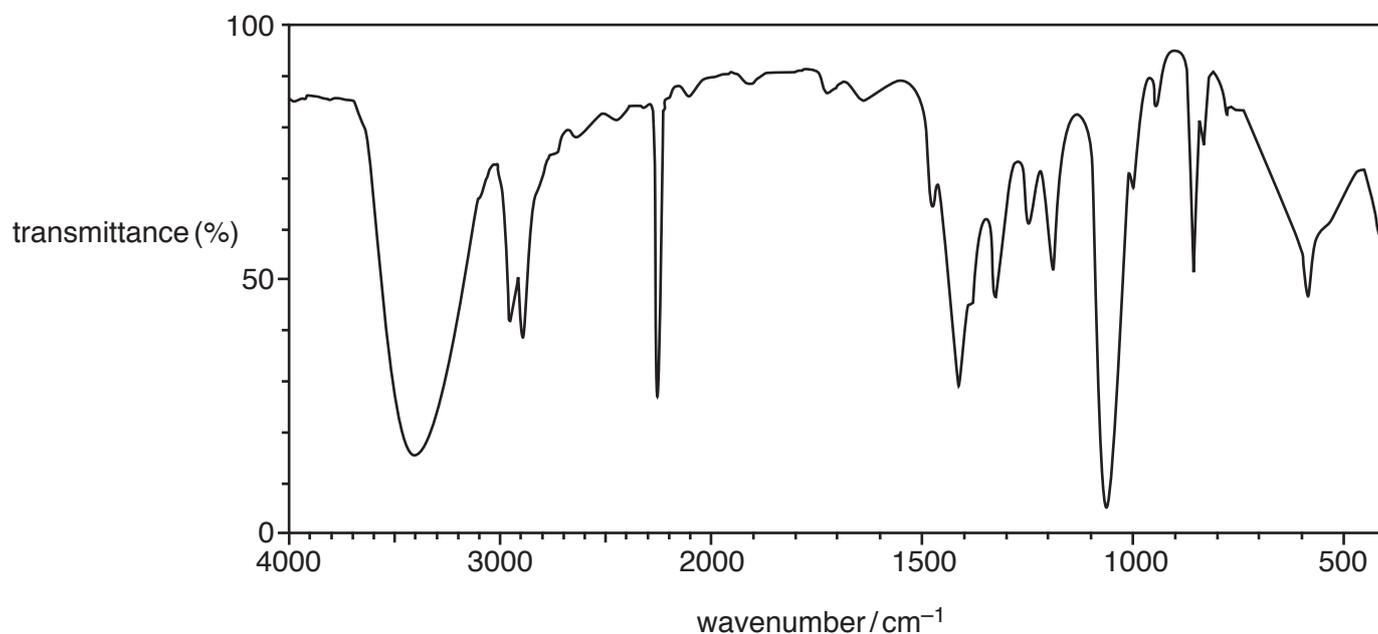
Which equation is a propagation step in the mechanism for this process?

- A  $\text{NO} + \text{O}_2 \rightarrow \text{N} + \text{O}_3$   
 B  $\text{NO} + \text{O}_2 \rightarrow \text{NO}_2 + \text{O}$   
 C  $\text{N} + \text{O}_3 \rightarrow \text{NO} + \text{O}_2$   
 D  $\text{NO}_2 + \text{O} \rightarrow \text{NO} + \text{O}_2$

Your answer

[1]

9 Which compound could have produced the IR spectrum below?



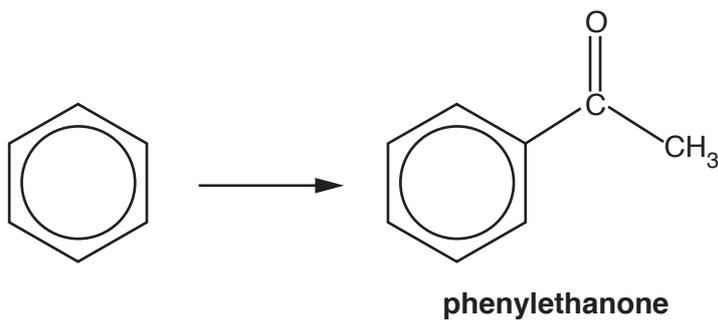
- A  $\text{CH}_3\text{CH}_2\text{OH}$   
 B  $\text{CH}_3\text{CHOHCN}$   
 C  $\text{CH}_3\text{COOH}$   
 D  $\text{CH}_3\text{CONH}_2$

Your answer

[1]

6

- 10 Benzene reacts with an organic reagent in the presence of a halogen carrier to form phenylethanone.



Which organic reagent is required?

- A  $\text{CH}_3\text{CH}_2\text{OH}$
- B  $\text{CH}_3\text{CHO}$
- C  $\text{CH}_3\text{COCl}$
- D  $\text{CH}_3\text{COOH}$

Your answer

[1]

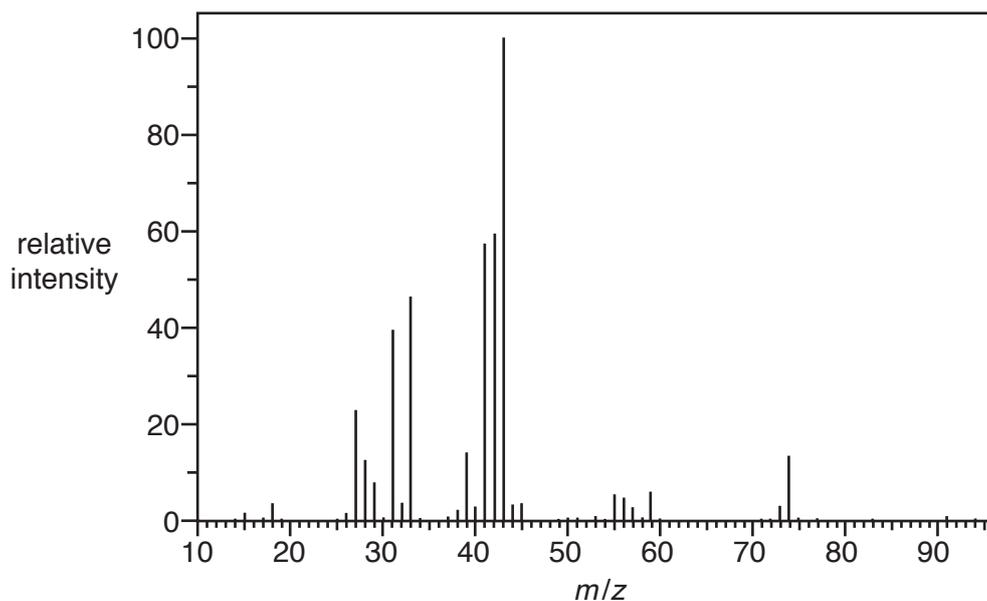
- 11 How many straight-chain structural isomers of  $\text{C}_7\text{H}_{15}\text{Cl}$  contain a chiral carbon atom?

- A 1
- B 2
- C 3
- D 4

Your answer

[1]

12 The mass spectrum of  $(\text{CH}_3)_2\text{CHCH}_2\text{OH}$  is shown below.



Which ion is responsible for the peak with the greatest relative intensity?

- A  $\text{CHCH}_2\text{OH}^+$
- B  $\text{CH}_3\text{CH}_2\text{CH}^+$
- C  $(\text{CH}_3)_2\text{CH}^+$
- D  $\text{CH}_3\text{CO}^+$

Your answer

[1]

13 Which statement(s) support(s) the delocalised model for the structure of benzene?

- 1 All carbon-carbon bonds have the same length.
- 2 The enthalpy change of hydrogenation of benzene is less exothermic than expected.
- 3 Bromine reacts with benzene less readily than with cyclohexene.

- A 1, 2 and 3
- B Only 1 and 2
- C Only 2 and 3
- D Only 1

Your answer

[1]

14 A solid organic compound can be purified by recrystallisation.

Which statement(s) about recrystallisation is/are true?

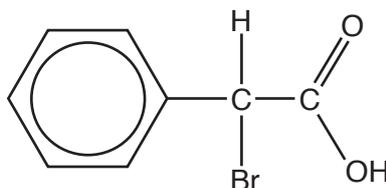
- 1 The organic compound is more soluble in hot solvent.
- 2 The hot solution is cooled before the purified organic compound is collected.
- 3 The melting point of the purified organic compound is lower than the impure compound.

- A** 1, 2 and 3  
**B** Only 1 and 2  
**C** Only 2 and 3  
**D** Only 1

Your answer

[1]

15 Which of the following could react with the compound below to form a carbon–carbon bond?



- 1  $\text{CH}_3\text{Cl}$  and  $\text{AlCl}_3$
  - 2 KCN in ethanol
  - 3  $\text{CH}_3\text{OH}$  and  $\text{H}_2\text{SO}_4$
- A** 1, 2 and 3  
**B** Only 1 and 2  
**C** Only 2 and 3  
**D** Only 1

Your answer

[1]

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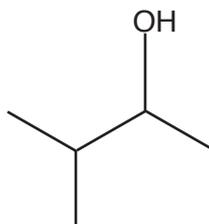
10

## SECTION B

Answer **all** the questions.

16 This question is about reactions of organic compounds containing carbon, hydrogen and oxygen.

(a) A chemist investigates two reactions of alcohol **A**, shown below.



**alcohol A**

(i) What is the systematic name of alcohol **A**?

..... [1]

(ii) What is the structural formula of alcohol **A**?

..... [1]

(iii) The chemist heats alcohol **A** with an acid catalyst to form a mixture containing **two** alkenes.

Draw the structures of the **two** alkenes formed in this reaction.

--	--

[2]

(iv) The chemist heats alcohol **A** with sodium chloride and sulfuric acid.

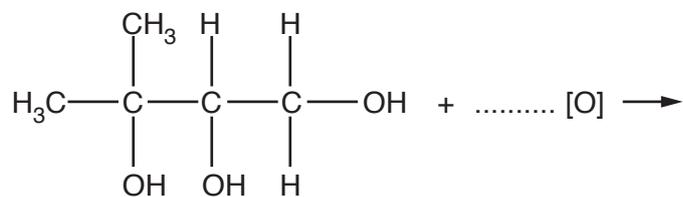
Construct a balanced equation for this reaction.

Show structures for the organic compounds in your equation.

[2]

- (b) Compound **B**, shown below, is refluxed with excess acidified potassium dichromate(VI) to form a single organic product.

Complete the equation for this reaction.

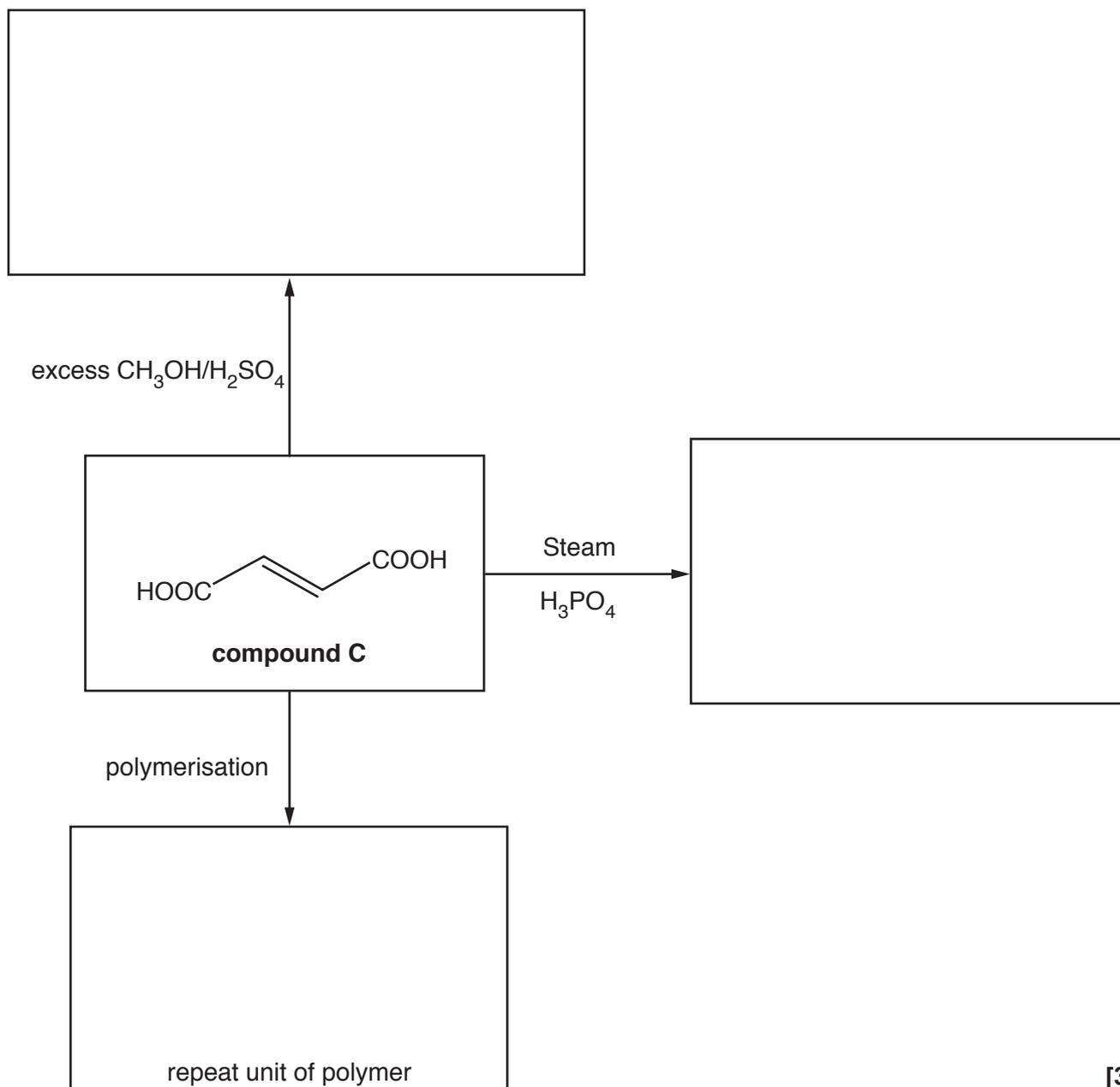


**compound B**

[2]

- (c) The flowchart below shows some reactions of compound **C**.

In the boxes, draw the organic products of these reactions.



[3]

Turn over

12

17 The general formula of an  $\alpha$ -amino acid is  $\text{RCH}(\text{NH}_2)\text{COOH}$ .

(a) The  $\alpha$ -amino acid cysteine ( $\text{R} = \text{CH}_2\text{SH}$ ) shows optical isomerism.

Draw 3-D diagrams to show the optical isomers of cysteine.

[2]

(b) The  $\alpha$ -amino acid lysine ( $\text{R} = (\text{CH}_2)_4\text{NH}_2$ ) reacts with an excess of dilute hydrochloric acid to form a salt.

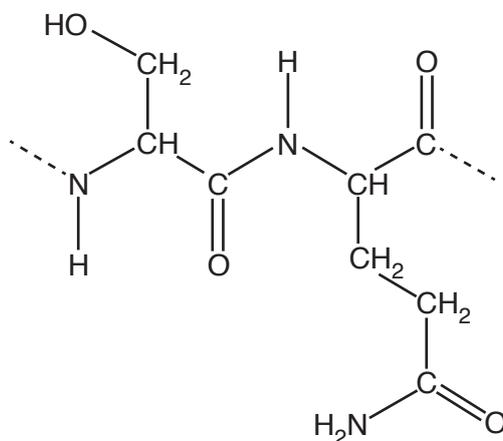
Draw the structure of the salt formed in this reaction.

[2]

13

(c)  $\alpha$ -Amino acids can react to form proteins.

A short section of a protein chain is shown below.



A student hydrolyses the protein with hot  $\text{NaOH}(\text{aq})$ .

Draw the structures of the organic products formed from this section of the protein.

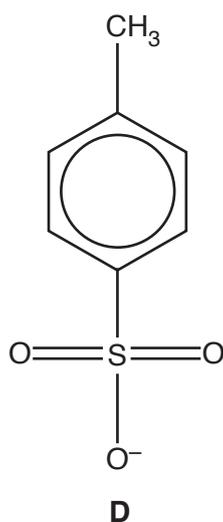
[3]





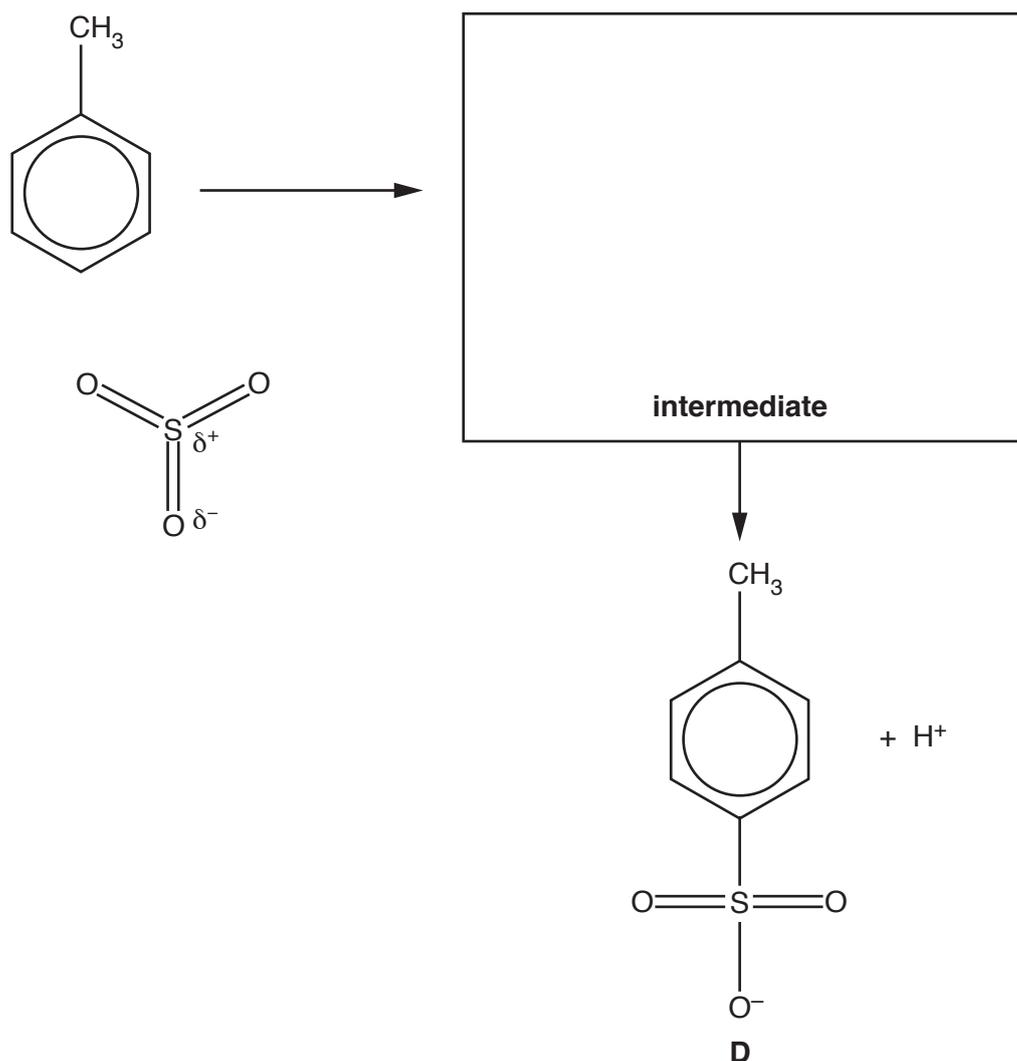
16

(b) Methylbenzene reacts with sulfur trioxide,  $\text{SO}_3$ , to form **D**, shown below.



The electrophile in this reaction is  $\text{SO}_3$ .

Complete the mechanism for the formation of **D**.  
Show curly arrows and the structure of the intermediate.



19 This question is about the hydrolysis of haloalkanes.

(a) The rate of hydrolysis of a haloalkane depends on the halogen present.

State and explain how the halogen in the haloalkane affects the rate of hydrolysis.

.....

.....

.....

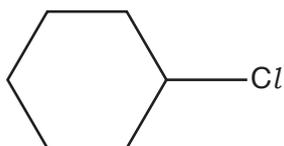
.....

..... [2]

(b) Chlorocyclohexane is hydrolysed with aqueous sodium hydroxide.

Outline the mechanism for this reaction.

Show curly arrows, relevant dipoles and the products.



[3]

(c) A student hydrolyses a haloalkane, **E**, using the following method.

- 0.0100 mol of haloalkane **E** is refluxed with excess NaOH(aq) to form a reaction mixture containing an organic product **F**.
- The reaction mixture is neutralised with dilute nitric acid.
- Excess AgNO<sub>3</sub>(aq) is added to the reaction mixture. 1.88 g of a precipitate **G** forms.

Organic product, **F**, has a molar mass of 74.0 g mol<sup>-1</sup> and has a chiral carbon atom.

- (i) Draw a **labelled** diagram to show how the student would carry out the hydrolysis of haloalkane **E**.

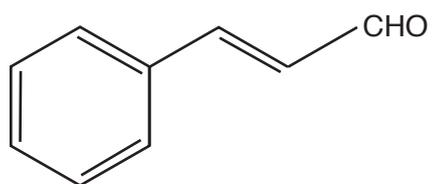
[2]

- (ii) Analyse the information to identify **E**, **F** and **G**.

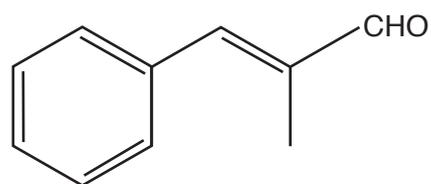
Show your working.

[3]

20 Cinnamaldehyde and methylcinnamaldehyde are naturally occurring organic compounds.



**cinnamaldehyde**



**methylcinnamaldehyde**

(a) Methylcinnamaldehyde is an *E* stereoisomer.

Explain this statement in terms of the Cahn-Ingold-Prelog (CIP) rules.

.....

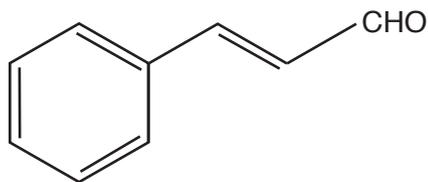
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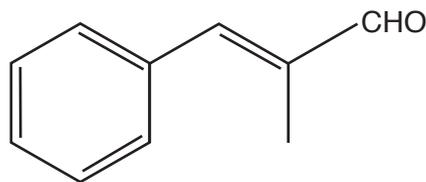
.....

..... [2]

- (b) A student plans to carry out some chemical tests on both cinnamaldehyde and methylcinnamaldehyde.



cinnamaldehyde



methylcinnamaldehyde

- (i) Suggest a suitable chemical test to confirm that both compounds contain an unsaturated carbon chain.

Your answer should include the reagent and observations.

.....  
 ..... [1]

- (ii) Describe a chemical test to confirm that both compounds contain an aldehyde functional group.

Your answer should include the reagent and observations.

.....  
 ..... [1]

- (iii) Describe a chemical test to confirm that cinnamaldehyde and methylcinnamaldehyde contain a carbonyl group.

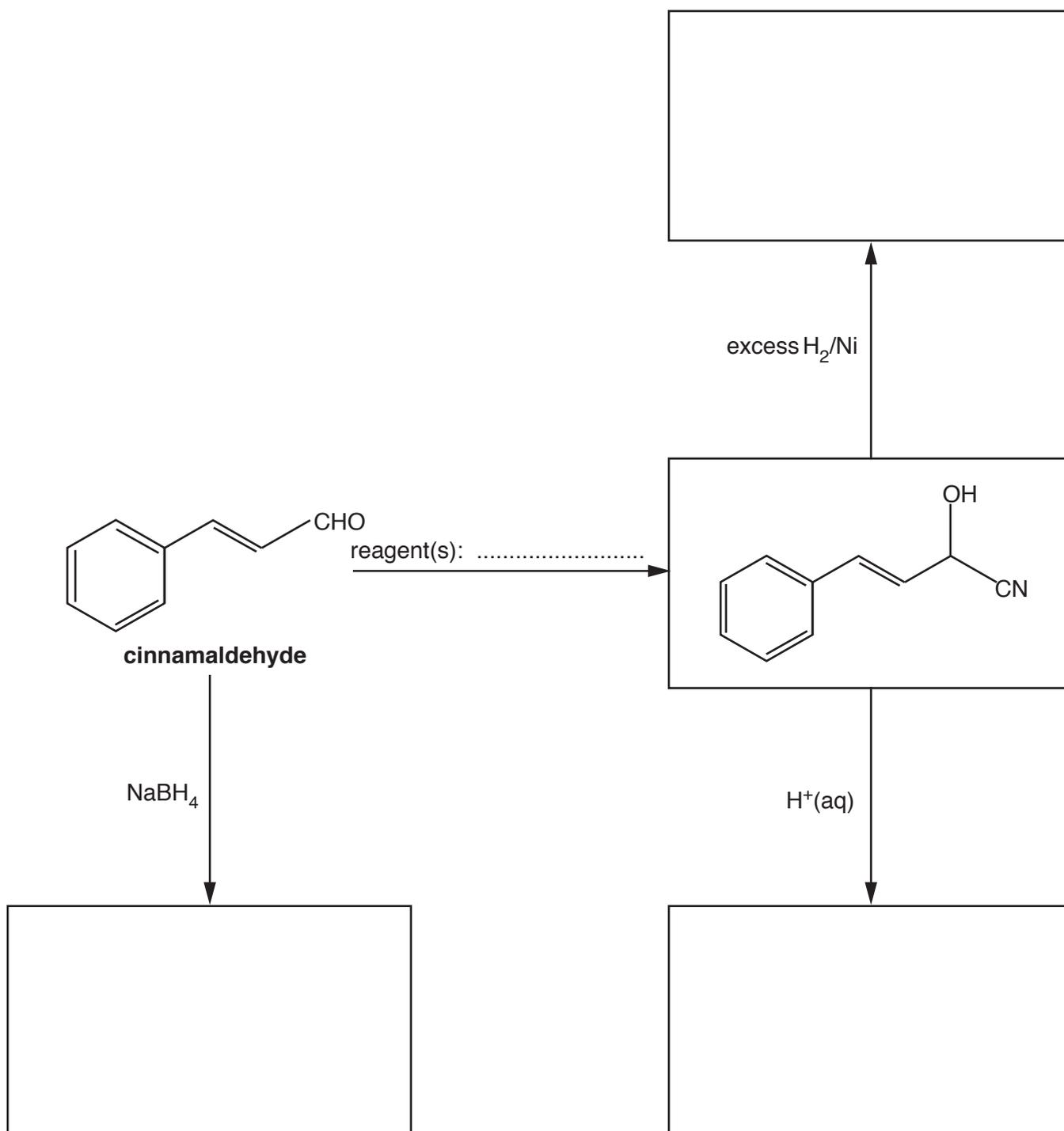
How could the products of this test be used to distinguish between the two compounds?

Your answer should **not** include spectroscopy.

.....  
 .....  
 .....  
 .....  
 .....  
 .....  
 ..... [3]

(c) The flowchart below shows some reactions starting with cinnamaldehyde.

Draw the structures of the missing organic compounds in the boxes and add the missing reagent(s) on the dotted line.



[5]



Additional answer space if required.

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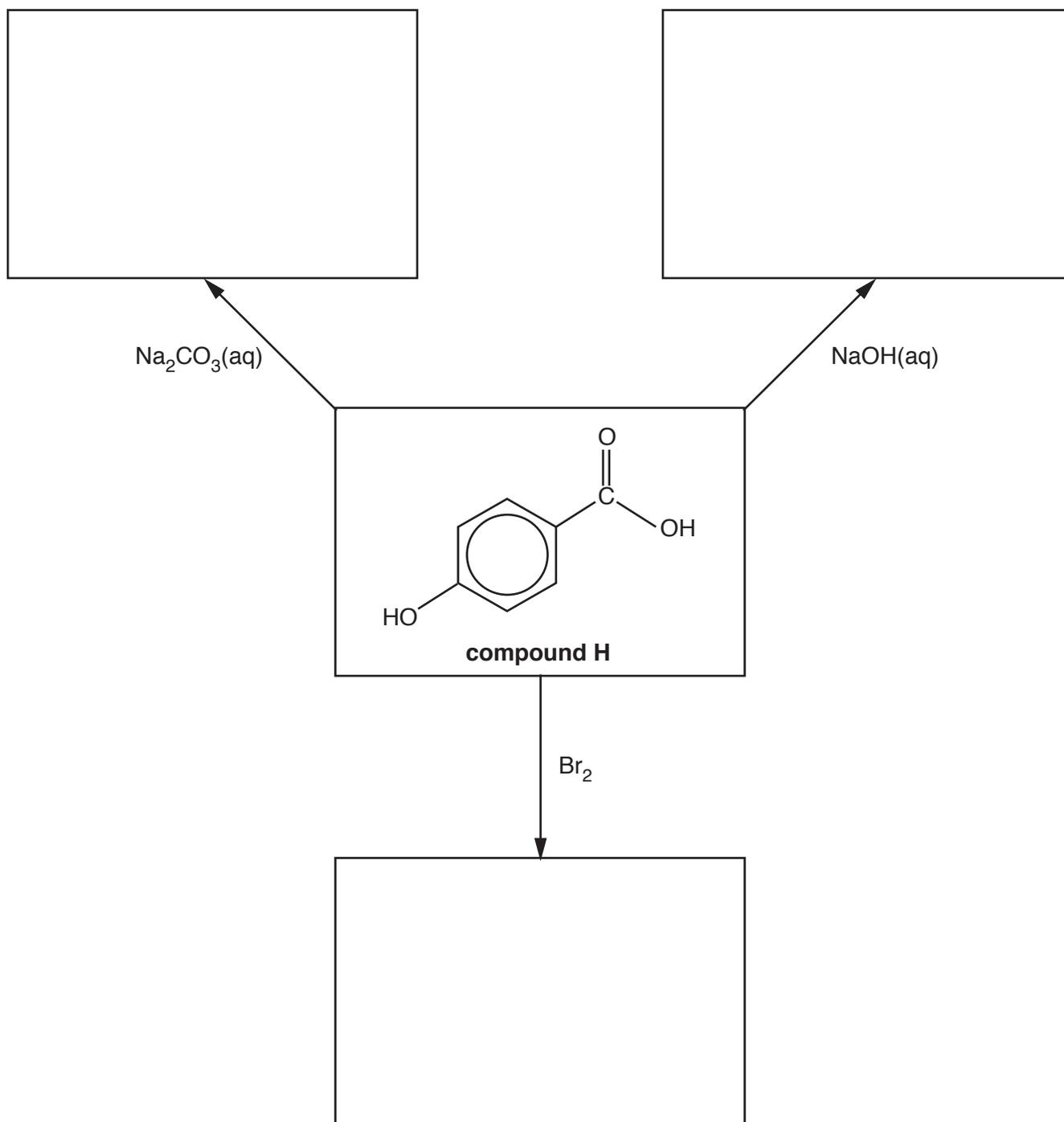
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21 This question is about aromatic carboxylic acids and their derivatives.

(a) The flowchart below shows some reactions of compound **H**.

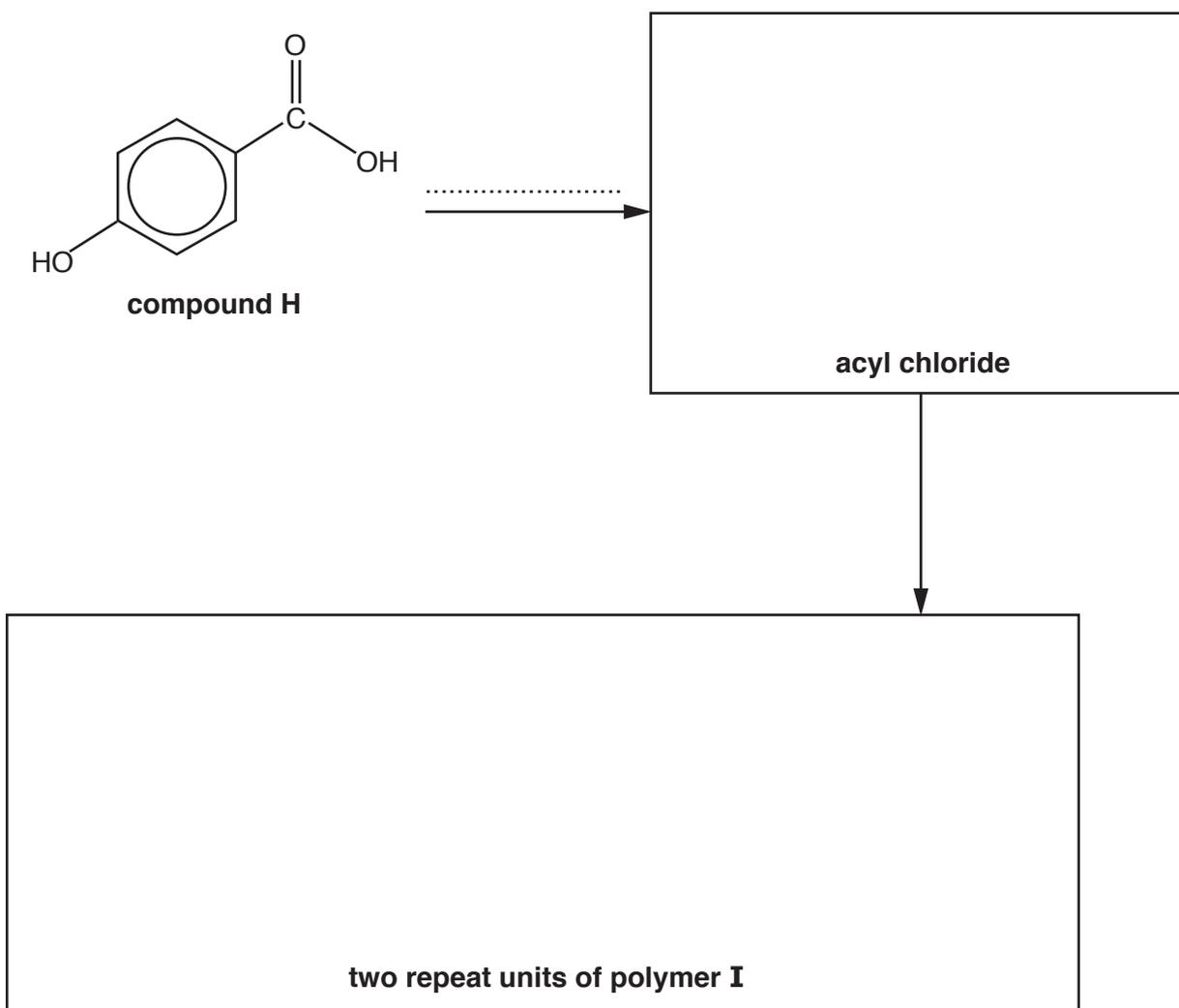
In the boxes, draw the organic products of these reactions.



[3]

(b) Compound **H** is used in the synthesis of polymer **I**, as shown in the flowchart below.

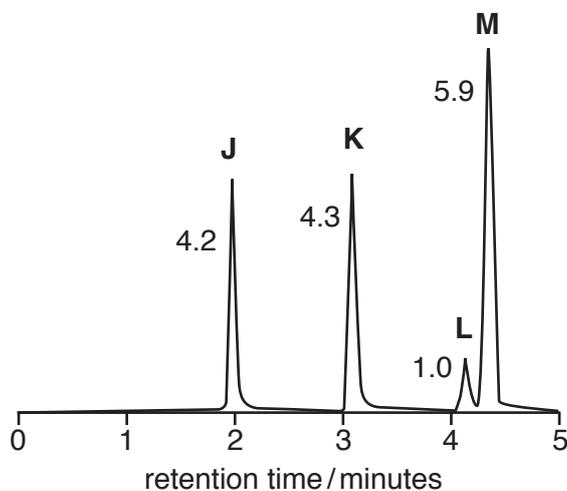
Complete the flowchart by drawing the structure of the acyl chloride and **two** repeat units of polymer **I**, and stating the **formula** of the reagent(s) required for the first stage on the dotted line.



[4]

- (c) A cosmetic product containing four esters, **J**, **K**, **L** and **M**, is analysed by gas chromatography and mass spectrometry. The results are shown below.

### Gas chromatogram



The numbers by the peaks are the relative molar proportions of the compounds in the mixture.

### Mass spectrometry

ester	<i>m/z</i> of molecular ion peak
<b>J</b>	152
<b>K</b>	166
<b>L</b>	180
<b>M</b>	180

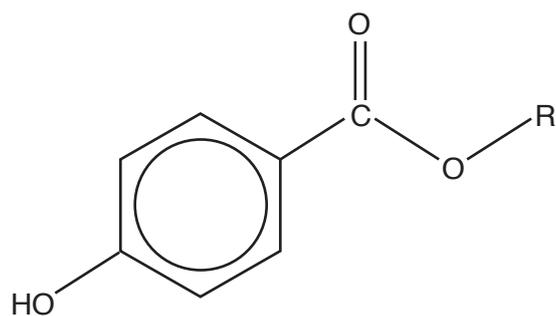
- (i) The concentration of ester **K** in the cosmetic product is  $9.13 \times 10^{-2} \text{ g dm}^{-3}$ .

Using the results, calculate the concentration, in  $\text{mol dm}^{-3}$ , of ester **M** in the cosmetic product.

Give your answer to **two** significant figures.

concentration of ester **M** = .....  $\text{mol dm}^{-3}$  [2]

(ii) A general structure for esters **J**, **L** and **M** is shown below.



Where 'R' is an alkyl group.

Use the mass spectrometry results to deduce possible structures for esters **J**, **L** and **M**.

<p style="text-align: center;"><b>J</b></p>	<p style="text-align: center;"><b>L</b></p>	<p style="text-align: center;"><b>M</b></p>
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[3]

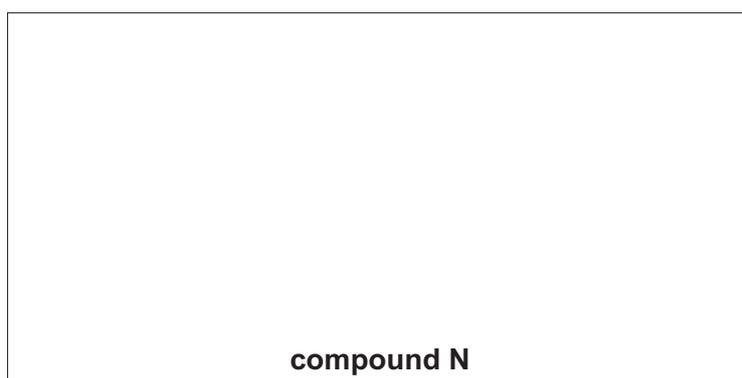


(c) Fuel additives are often used to improve the combustion of a fuel.

(i) Compound **N** is a fuel additive containing carbon, hydrogen and oxygen only.

Complete combustion of 1.71 g of compound **N** produces 2.97 g of  $\text{CO}_2$  and 1.62 g of  $\text{H}_2\text{O}$ . The relative molecular mass of compound **N** is 76.0.

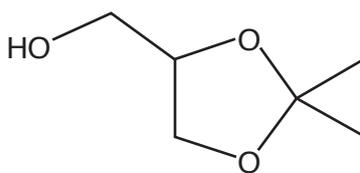
Calculate the molecular formula of **N** and suggest a possible structure for the compound.



[5]

30

(ii) Solketal has been investigated as a potential fuel additive.



solketal

Solketal is synthesised from propane-1,2,3-triol and a carbonyl compound.

Construct a balanced equation for this synthesis.

Show structures for the organic compounds in your equation.

[2]

(d)\* A scientist is researching compounds that might be suitable as fuel additives. One of the compounds gives the analytical results below.

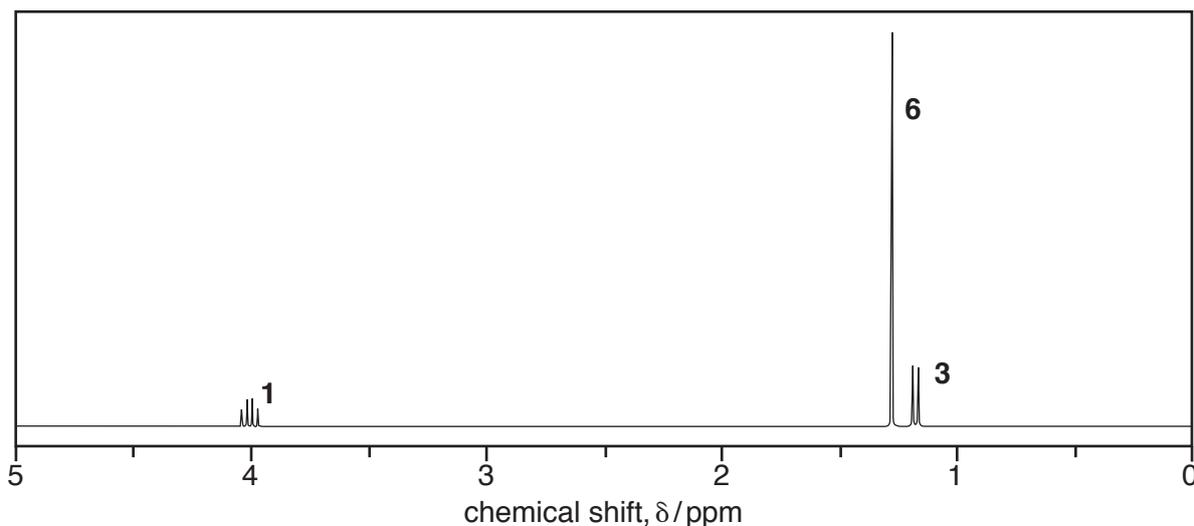
**Elemental analysis by mass:**

C: 54.54%; H: 9.10%; O: 36.36%

**Mass spectrum:**

Molecular ion peak at  $m/z = 132.0$

**<sup>1</sup>H NMR spectrum in D<sub>2</sub>O**



The numbers by the peaks are the relative peak areas.

When the spectrum is run without D<sub>2</sub>O, there are **two** additional peaks with the same relative peak areas at 11.0 ppm and 3.6 ppm.



**ADDITIONAL ANSWER SPACE**

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

A large rectangular area with a solid vertical line on the left side and horizontal dotted lines across the rest of the page, providing space for writing answers.



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