



GCE

Chemistry B (Salters)

Unit **H033/01**: Foundations of chemistry

Advanced Subsidiary GCE

Mark Scheme for June 2016

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

OCR will not enter into any discussion or correspondence in connection with this mark scheme.













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H033/01

Mark Scheme

June 2016

Annotations

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Benefit of doubt not given
	Noted but no credit given
	Ignore
	Blank page

Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
✓	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

H033/01

Mark Scheme

June 2016

SECTION A

Question	Answer	Marks	AO element
1	C	1	1.1
2	B	1	1.1
3	C	1	1.2
4	B	1	2.5
5	C	1	1.1
6	A	1	1.1
7	B	1	1.1
8	D	1	1.2
9	B	1	1.2
10	D	1	2.5
11	A	1	2.4
12	D	1	2.5
13	C	1	2.1
14	D	1	2.4
15	D	1	2.1
16	C	1	1.2
17	D	1	2.7
18	D	1	2.2
19	C	1	2.4
20	B	1	1.1

H033/01

Mark Scheme

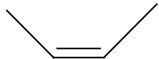
June 2016

Question			Answer	Marks	AO element	Guidance
21	(a)		mass number 56 protons 28 neutrons 28	1	2.2	
	(b)		$(28 \times 92.17) + (29 \times 4.71) + (30 \times 3.12)$ /100 = 28.1(095) ✓ 28.11 (2dp) ✓	2	1.2 1.2	NO ecf 28.11 alone on the answer line scores 1 mark without some evidence of working. Answers to other dp without working score zero.
	(c)	(i)	1. (Add NiCO_3 to H_2SO_4) until fizzing/reaction stops/all sulfuric acid has reacted ✓ 2. filter (prior to crystallisation) ✓ 3. (partially) <u>evaporate</u> ✓ 4. Filter the crystals/pick out the crystals/leave crystals to dry/fully evaporate to produce crystals ✓	4	1.2 1.2 1.2 1.2	2. IGNORE what is being filtered but NOT nickel sulfate 3. ALLOW a word derived from evaporate e.g. evaporation 4. boiling to dryness CON pt 4, ALLOW products/ $\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$ for "crystals" fully evaporate to produce crystals scores pts 3 & 4
		(ii)	A no – this would have resulted in a higher mass/yield ✓ B no – nickel carbonate is added in <u>excess</u> AW/ moles of sulphuric acid are the limiting factor/(excess)nickel carbonate is removed ✓ C yes – loss of water would reduce <u>mass</u> (AW) ✓	3	3.1 3.1 3.1	for each mark, 'yes' or 'no' (or 'correct'/'incorrect') must be stated (or implied) and the reason given. ALLOW weight for mass
				10		

H033/01

Mark Scheme

June 2016

Question			Answer	Marks	AO element	Guidance
22	(a)		(2-)methylpropene	1	1.2	IGNORE dashes, commas and gaps ALLOWprop-1-ene ALLOW minor spelling errors
	(b)			1	2.5	ALLOW any unambiguous representation of the structure but NOT C for CH ₃ IGNORE name
	(c)	(i)	water/steam AND phosphoric/ concentrated sulfuric acid	1	1.2	ALLOW names of acids (including oxidation states) or formulae (c.H ₂ SO ₄ , H ₂ SO ₄ (l), H ₃ PO ₄) but if name AND formula present both must be correct. IGNORE isobutylene/2-methylpropene
	(c)	(ii)	tertiary AND OH/functional group attached/bonded to C that has... no H atoms or 3 C/methyl/alkyl/R (groups)	1	1.1	IGNORE 'it' for 'OH' ALLOW Alcohol/hydroxyl group attached to C..... NOT hydroxide
		(iii)	(heat with) acidified dichromate(VI) ✓ A/tertiary has no reaction/stays orange /doesn't change colour. Others/primary/secondary go green/change colour ✓	2	3.4 3.4	Mark separately ALLOW correct formulae ALLOW sulfuric acid or sulfuric(VI) acid as replacement for acidified. IGNORE formulae if names correct ALLOW dichromate without 'VI' as long as no other number is there
		(iv)	forms the weakest/smallest/ fewest instantaneous (dipole) – induced dipole bonds/forces ✓ smallest surface area OR non-linear molecules are unable to get closer together/align AW ✓	2	2.1 1.2	Mark separately IGNORE references to other intermolecular bonding ALLOW weaker/smaller/fewer ALLOW Van der Waal's forces or London forces DO NOT ALLOW id-id in first instance ALLOW minor spelling errors ALLOW 'less contact between molecules'
	(d)	(i)	Ether/alkoxy/methoxy	1	1.2	
		(ii)	C ₄ H ₈ + CH ₄ O → C ₅ H ₁₂ O	1	2.6	must be molecular formulae as shown ALLOW element symbols in any order IGNORE state symbols
				10		

H033/01

Mark Scheme

June 2016

Question			Answer	Marks	AO element	Guidance
23	(a)	(i)	$\text{Cl}_2 + 2\text{Br}^- \rightarrow 2\text{Cl}^- + \text{Br}_2$	1	1.2	IGNORE state symbols
		(ii)	chlorine is more reactive and gains electrons more readily (than bromine) <i>ora</i> OR and removes electrons from bromide/bromine (ions)	1	1.1	IGNORE other references to halides IGNORE references to electronegativity IGNORE "attract" ALLOW "accept"
	(b)	(i)	move through: $\text{Na}^+ \checkmark$ not through: Cl^- and $\text{OH}^- \checkmark$	2	3.1 3.1	IGNORE H^+ any other ions are CON in each case ALLOW names of ions
		(ii)	$2\text{NaCl} + 2\text{H}_2\text{O} \rightarrow \text{Cl}_2 + \text{H}_2 + 2\text{NaOH}$	1	2.6	ALLOW multiples or halves IGNORE state symbols
		(iii)	100% since... no waste or all useful/desired products	1	3.2	IGNORE by-products
		(iv)	1. 37000/40 OR 925 \checkmark 2. 925x24/2 OR 11100 \checkmark 3. $1.1 \times 10^4 \text{ (dm}^3\text{)} \checkmark$ (2 sf and standard form)	3	2.2 2.2 2.2	Allow ecf from 1. and 2. No ecf from (b)(ii) without reference to working: 1.1 x 10 ⁴ scores 3 marks 2.2 x 10 ⁴ scores 2 marks 11100 or 11000 score 2 marks 22000 or 22200 score 1 mark 11 or 1.1 x 10 ¹ scores 2 marks 22 or 2.2 x 10 ¹ scores 1 mark 9.3 x 10 ² scores 2 marks 0.93 scores 1 mark
	(c)	(i)	-1, +5, 0	1	2.4	NOT signs after numbers ALLOW roman numerals (-I; +V)
		(ii)	(Br in) BrO_3^- because oxidation state (of Br) goes down (+5 to 0)	1	1.2	ALLOW 'bromate' or 'bromate(V)' for BrO_3^- ALLOW 'bromine reduced' IF correct oxidation states given ALLOW ecf on oxidation states from (c)(i)
				11		

H033/01

Mark Scheme

June 2016

Question			Answer	Marks	AO element	Guidance
24	(a)	(i)	atom/molecule/species/ ion with <u>unpaired</u> electron(s) ✓ formed from chloroalkanes/ R-Cl/ haloalkanes/ chlorine compounds/ CFCs/ C-Cl ✓ by high-energy/ high frequency ✓ uv ✓ homolytic (fission)✓	5	1.1 1.1 1.1 1.1 1.1	IGNORE 'single' 'lone' ALLOW 'it' for 'atom/molecule/species/ ion' ALLOW equation showing formation of radicals Mention of breakdown of Cl ₂ CONS 2nd marking point IGNORE UVA for third marking point UVB and/or C covers both 3rd and 4th marking points. ALLOW homolysis
		(ii)	Cl + O ₃ ----> ClO + O ₂ ClO + O ----> Cl + O ₂ ✓ O ₃ + O ----> 2O ₂ ✓	2	1.2 1.2	Mark separately IGNORE dots Cl + O ₃ ----> ClO + O ₂ ClO + O ₃ ----> Cl + 2 O ₂ and hence 2O ₃ ----> 3O ₂ scores 2 nd marking point only
	(b)	(i)	Y axis (concentration) correctly labelled with units and standard form ✓ X axis (time) correctly labelled with units ✓ Points plotted utilising over half of each axis ✓	3	2.6 2.6 2.6	Linear scale labelled with conc(entration) of ozone/O ₃ NOT [O ₃] AND units (molecules cm ⁻³). 10 ¹² or 10 ⁻¹² mentioned somewhere. Linear scale labelled with "Time" AND units. X and y axes swapped but otherwise correct scores only one of the first two marks. Exact position of points does not need checking. Line of best fit not essential but point to point or curved line is CON
		(ii)	4.979 – 4.983 x 10 ¹² molecules cm ⁻³	1	2.6	Minimum of 4 sf required
		(iii)	it remains constant (AW) AND the gradient is constant/straight line graph (AW)	1	2.7	NOT almost/fairly/nearly constant IGNORE references to negative rate

Question			Answer	Marks	AO element	Guidance
	(c)		Rearrangement of gas equation to: $n = PV/RT$ ✓	4	2.5	can be implied by later working
			Conversion of volume units to m^3 ✓		2.6	can be implied from working
			$n = P \times \text{evaluated volume} / 8.314 \times 300$ ✓		2.6	ALLOW ecf if gas equation has been incorrectly rearranged wrong evaluation of this expression is CON
			Conversion of moles to molecules and evaluate to 2 or more sf. (correct answer is $2.4(1) \times 10^{17}$) ✓		2.6	Correct answer alone scores 4 marks <i>with no working:</i> 2.4×10^{23} or 2.4×10^{20} (incorrect conversion of v) scores 3 marks 4.0×10^{-7} (no N_A) scores 3 marks 0.4..... (incorrect v and no N_A) scores 2 marks

H033/01

Mark Scheme

June 2016

Question		Answer	Marks	AO element	Guidance
	(d)	<p>Choice of method:</p> <p>EITHER <i>Calculate the energy of 1 mole of photons and compare with bond enthalpy.</i> Calculate energy of one photon ($9.5 \times 10^{14} \times 6.63 \times 10^{-34}$ OR 6.30×10^{-19}) ✓ Multiply up for 1 mole and convert to kJ (photon energy x $6.02 \times 10^{23}/1000$) ✓ Evaluate (379) and state whether bond will be broken. ✓</p> <p>OR <i>Compare the energy of one photon with one bond</i> Calculate energy of one photon ($9.5 \times 10^{14} \times 6.63 \times 10^{-34}$ OR 6.30×10^{-19}) ✓ Calculate energy of one bond ($302 \times 1000/6.02 \times 10^{23}$) ✓ Evaluate both (6.30×10^{-19} and 5.02×10^{-19}) and state whether bond will be broken. ✓</p> <p>OR <i>Calculate minimum frequency needed to break bond</i> Calculate energy required per molecule ($302000/6.02 \times 10^{23}$) ✓ Calculate required frequency of radiation (energy/6.63×10^{-34}) ✓ Evaluate (7.57×10^{14}) and state whether bond will be broken. ✓</p>	3		The 3 rd marking point can only be scored if the first 2 marks are scored OR if the only error is in conversion between J and kJ i.e. failure to convert or incorrect conversion
				3.2	
				3.2	
				3.2	
				3.2	
				3.2	
				3.2	
				3.2	
				3.2	
				19	

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